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produce carbon disulfide and hydrogen gas.  $\text{CH}_4(\text{g}) + 2\text{H}_2\text{S}(\text{g}) \rightleftharpoons \text{CS}_2(\text{g}) + 4\text{H}_2(\text{g})$ ;  
 $\Delta H = +232.5 \text{ kJ} \dots$

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$10^{-3}$  and  $10^3$ . Chemical equilibrium (definition) The state reached when the concentration of

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reactants and products remain constant over time. What is the general, reversible reaction for equilibrium? (With letters)  $aA + bB \rightleftharpoons cC + dD$ . Equilibrium equation (general with letters)  $K_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}$ .

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26 D5 Chemical equilibrium is said to be dynamic because A. the reaction proceeds quickly. B. the mass of the reactants is decreasing. C. the macroscopic properties are constant. D. both forward and reverse reactions are occurring. 27 D5 Equilibrium is a dynamic process because the A. macroscopic properties are not changing.

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~~This model should show the three states that occur in the chemical equilibrium process. As students work, walk around to offer suggestions and guide. Have students:~~

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~~Chemical Equilibrium Study Guide Answers the stress. 2. Answers may vary, but should include changes in concentration, volume, and temperature. Some students may list pressure~~

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in place of volume. 3. left; remove N<sub>2</sub> or H<sub>2</sub> VIBRATIONS AND WAVES - Weebly Acces PDF Chapter 18 Chemical Equilibrium Study Guide Answers Chapter 18 Chemical Equilibrium Study Page 10/16

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Chemical Equilibrium Constant: A reversible chemical reaction system can proceed in both directions from a set of initial conditions (concentrations) towards a state of chemical equilibrium.

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The given reaction equation is:  $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$   $\Delta H = -191 \text{ kJ}$   
 $2\text{SO}_3(\text{g}) \rightleftharpoons 2\text{SO}_2(\text{g}) + \text{O}_2(\text{g})$   $\Delta H = +191 \text{ kJ}$ . The equilibrium constant is a ratio of the forward and reverse ...

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1) Give a balanced chemical equation which describes this equilibrium. 2) Explain what influence the removal of  $\text{H}_2$  from the container will have on the equilibrium. 3) Explain what effect...

~~You are given a container in which  $\text{H}_2(\text{g})$  ... study.com~~

Solved: Consider the following equilibrium:  $4\text{HCl}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g}) + 2\text{Cl}_2(\text{g})$ ;  $K_c = 357$ . An equilibrium mixture of the system was found to...

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Changing the temperature of a reaction at equilibrium alters both the equilibrium constant and the equilibrium position. When a reaction is exothermic, which means it releases energy, lowering the temperature shifts the equilibrium to the right because the forward reaction liberates heat and removes the reactants.

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