

Chapter 14 Chemical Periodicity Answers

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Chapter 14 Chemical Periodicity Answers

CHAPTER NOTES – CHAPTER 14 Chemical Periodicity Goals : To gain an understanding of : 1. Electron configurations 2. Periodicity. The periodic law states that when the elements are arranged according to increasing atomic number there is a periodic pattern in their physical and chemical properties.

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Trends in Atomic Size Influenced by three factors: 1. Energy Level Higher energy level is further away. 2. Charge on nucleus More charge pulls electrons in closer. 3. Shielding effect Group trends As we go down a group... each atom has another energy level, so the atoms get bigger. Periodic Trends As you go across a period, the radius gets smaller.

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with a term, short phrase, or number. The periodic table organizes the elements into vertical groups and horizontal ...

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Chapter 14 – Chemical Periodicity and Trends Watch the Professor Dave Explains video, The Periodic Table: Atomic Radius, Ionization Energy, and Electronegativity at <https://www.youtube.com/watch?v=hePb00CqvP0> Arrangement of the Periodic Table 1. Why was Dimitri Mendeleev ' s model for the period table more powerful than other models

Chapter 14 Chemical Periodicity and Trends

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7. Give the names and chemical symbols for the elements that correspond to these atomic numbers: a. 10 b. 18 c. 36 d. 90 8. List, by number, both the period and group of each of these elements. Symbol Period Group a. beryllium Be b. iron Fe c. lead Pb 9. Which of the following pairs of elements belong to the same period?

Chemistry: The Periodic Table and Periodicity

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Answer. 1. Representative elements: The Representative Elements are those elements within the first two families (Groups I and II on the far left) and the last six families or groups (on the right) of the Periodic Table. 2. d-transition elements: Set of metallic elements occupying a central block (Groups IVB – VIII, IB, and IIB, or 4 – 12) are d-transition elements.

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