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5. ACID, BASES AND SALT QUESTIONS AND ANSWERS/EXERCISE - SCIENCE CLASS 9 CHAPTER 5

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Chapter 5 -**Part 1 | Acids , Bases And Salts | PYQ Deep Discussion | Chemistry | Smart Ideas**

Acids Bases And Salts Answer

Solution: The reaction of the acid + base gives a product of salt + water, which is considered as neutralization reaction.

Examples: $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

NCERT Solutions for class 10 chapter 2 Acids, bases and salt

A salt is an ionic compound that can be formed by the neutralization reaction of an acid and a base. The following table shows the patterns for the solubility of salts . To

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help you remember, think of SPANE for soluble salts: Sodium, Potassium, Ammonium, Nitrates and Ethanoates.

*Acid, Bases, Salts - IGCSE Chemistry
(solutions, examples ...*

Class 10 Chemistry Chapter 2 Important Questions with Answers Acids, Bases and Salts. ...

Acids, Bases and Salts Class 10 Important Questions with ...

Let's learn more about Acids Bases and Salts!
We will discuss the important properties of

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Acids, Bases and Salts and how they differ from each other. Simple examples and interesting demos are shown in the video! Try the Top 3 Questions at the end of the video and write your answers & doubts in the comments below!

Acids Bases and Salts

Answer-1. Post-Your-Explanation-1. 2. Bases react with. acids to produce salts and water. water to produce acids and salts. salts to produce acids and water. neither acids, salts, nor water. Answer-2.

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Acids Bases and Salts multiple choice questions and ...

Answer A reaction in which an acid and base react with each other to give a salt and water is termed as neutralization

reaction. For Example: (i) $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ (ii) $\text{HNO}_3 + \text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$ 15.

Give two important uses of washing soda and baking soda. Answer Two important uses of washing soda are:

NCERT Solutions for Class 10th Science: Ch 2 Acids, Bases ...

Phenolphthalein turns colourless in acidic

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solution and turns pink in basic solution. Methyl orange turns red in acidic solution and turns yellow in basic solution. Answer: Some indicators show a change in their odour in the presence of acids or bases. Such indicators are called olfactory indicators.

Class 10 Short Answer Type Questions-Acid, Base-Salts

Answer. (a) HCl is acid and NaOH is base whose combination forms the common salt. Its formula is NaCl (Sodium chloride). It is obtained from sea water. (b) Rock salt is the common name for the mineral "halite". Its

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chemical formula is NaCl . It may be white or light blue or yellow depending upon impurities present in it.

Acids Bases and Salts Chapter Wise Important Questions ...

A substance used as an acid-base indicator.

(6) 11. An atom or molecule with a positive or negative charge. (3) 13. The strength of a base. (10) 15. Common name for sodium hydroxide. (3) 17. Color of litmus paper in acids. (3) Title: Acids, Bases, Salts

Crossword Subject: Science Author: Chris Gunn

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Acids, Bases, Salts Crossword - science-teachers.com

chemical reaction that occurs when the H_3O^+ ions form an acid and react with the OH^- ions from a base to produce water molecules and a salt. salt. compound formed when negative ions from an acid combine with positive ions from a base. titration.

Chapter 22-Acids, Bases, and Salts Flashcards / Quizlet

The reaction between an acid and a base is

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called neutralisation. In this process, salt and water are formed with the evolution of heat. Acid + Base Salt + Water + (heat is evolved) Here is a reaction showing neutralisation. Question 6. Why do you rub baking soda or calamine solution when an ant bites us? Answer:

Acids, Bases and Salts Class 7 Extra Questions and Answers ...

Class 10 Science Acid, Bases and Salts Mind Map Acids. Produce $[H^+]$ in H_2O ; Sour taste; Turns blue litmus red; Acts as electrolyte in solution; Chemical Properties Of Acids Acids

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+ Metal ? Salt + H₂ e.g. H₂SO₄ + Zn ?
ZnSO₄ + H₂ Acids + Metal Carbonate/Metal
hydrogen Carbonate = Salt + CO₂ e.g NaHCO₃
+ HCl ? NaCl + H₂O + CO₂ Acids + Bases ?
Salt + water

*NCERT Solutions for Class 10 Science Chapter
2 Acids ...*

1) salts are the chemical compounds
consisting of the ionic assembly of cations
and anions and these are neutral. Sodium
chloride is a salt An el view the full answer
Previous question Next question

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Solved: What Are Salts, Electrolytes, Bases, And Acids? Be ...

When an acidic solution is mixed with a basic solution, they neutralize each other, forming salt and water. [B]. Neutral is the solution which does not change the colour of either red or blue litmus.

Class 7 Science Chapter 5 MCQ of Acids, Bases and Salts ...

Play this game to review Chemistry. Which of the following values would represent the pH of a strong base?

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Acids/Bases and Salts | Chemistry Quiz - Quizizz

When added to an acidic solution, it turns red and when added to a basic solution, it turns blue. It is available in the form of a solution, or in the form of strips of paper, known as litmus paper. Generally, it is available as red and blue litmus paper. Hope the knowledge provided regarding the Acids, Bases and Salts Class 7 Extra Questions and Answers has helped you to the possible extent.

Acids, Bases and Salts Class 7 Extra
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Questions and Answers ...

Students can refer these Acids, Bases and Salts Class 10 MCQs Questions with Answers and assess their preparation level. Acids, Bases and Salts Class 10 MCQs Questions with Answers. Question 1. A strong acid or base :
(a) is stable in solution (b) has a higher molecular weight (c) has a high boiling point (d) ionises completely in solution. Answer:

MCQ Questions for Class 10 Science Chapter 2 Acids, Bases ...

MCQ Questions for Class 7 Science with Answers were prepared based on the latest

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exam pattern. We have Provided Acids, Bases and Salts Class 7 Science MCQs Questions with Answers to help students understand the concept very well. Acids, Bases and Salts Class 7 MCQs Questions with Answers. Choose the correct answer: Question 1.

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